

# New energy battery grade classification diagram

How are batteries classified?

Batteries can be classified according to their chemistry or specific electrochemical composition, which heavily dictates the reactions that will occur within the cells to convert chemical to electrical energy. Battery chemistry tells the electrode and electrolyte materials to be used for the battery construction.

What are the different types of primary batteries?

Primary batteries come in three major chemistries: (1) zinc-carbon and (2) alkaline zinc-manganese, and (3) lithium (or lithium-metal) battery. Zinc-carbon batteries is among the earliest commercially available primary cells. It is composed of a solid, high-purity zinc anode (99.99%).

What are the different types of EV batteries?

Three main types of batteries dominate today's EV market: Lithium Iron Phosphate (LFP), Nickel Manganese Cobalt (NMC), and Nickel Cobalt Aluminum (NCA) batteries. According to the IEA's 2024 report, LFP and NMC batteries together account for over 90% of the global EV battery market.

What is a secondary battery chemistry?

Secondary battery chemistries, distinct from primary batteries, are rechargeable systems where the electrochemical reactions are reversible. Unlike primary batteries that are typically single-use, secondary batteries, such as lithium-ion and nickel-metal hydride, allow for repeated charging and discharging cycles.

What is battery chemistry?

Battery chemistry tells the electrode and electrolyte materials to be used for the battery construction. It influences the electrochemical performance, energy density, operating life, and applicability of the battery for different applications. Primary batteries are "dry cells".

What is the difference between alkaline and secondary battery chemistries?

An alkaline battery is capable of providing approximately three to five times the energy output compared to a zinc-carbon dry cell of equivalent size. Secondary battery chemistries, distinct from primary batteries, are rechargeable systems where the electrochemical reactions are reversible.

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